Energetic Comparison between Photoinduced Electron-Transfer Reactions from NADH Model Compounds to Organic and Inorganic Oxidants and Hydride-Transfer Reactions from NADH Model Compounds to *p*-Benzoquinone Derivatives

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Abstract: Kinetic studies on photoinduced electron-transfer reactions from dihydropyridine compounds (PyH_2) as being NADH model compounds to organic and inorganic oxidants and hydride-transfer reactions from PyH_2 to *p*-benzoquinone derivatives (Q) in the absence and presence of Mg^{2+} ion are reported by determining over 150 rate constants. These results, combined with the values of Gibbs energy change of the photoinduced electron-transfer reactions as well as those of each step of the hydride-transfer reactions as being the e^-–H^+–e^- sequence, which are determined independently, revealed that the rate constants of the photoinduced electron-transfer reactions obey the Rehm–Weller–Gibbs energy relationship and that the activation barrier of the hydride-transfer reactions from PyH_2 to Q is dependent solely on the Gibbs energy changes of the initial electron transfer from PyH_2 to Q and the following proton transfer from $PyH_2^{\cdot+}$ to $Q^{\cdot-}$ and thus independent of the Gibbs energy change of the final electron transfer from $PyH^{\cdot}$ to $QH^{\cdot}$. The retarding effect of Mg^{2+} ion observed on the photoinduced electron transfer and hydride-transfer reactions of PyH_2 is ascribed to the positive shifts of the redox potentials of the ground and excited states of PyH_2 due to the complex formation with Mg^{2+} ion.

Many redox reactions that had been viewed simply as direct transfer of a two-electron equivalent such as a hydride ion or electrophile–nucleophile binding have recently been reconsidered as proceeding via one-electron transfer in the activation process, followed by the reaction in a resulting radical ion pair.[1-4] Since in many cases, the products and kinetics are the same regardless of the mechanisms, it seems to be difficult to distinguish between these two mechanisms, often creating mechanistic controversies,[5] although in some cases detection of radical intermediates or isolation of products that could have arisen only via radicals represents good evidence for the electron-transfer mechanisms.[4] Especially, mechanisms of hydride-transfer reactions from reduced nicotinamide adenine dinucleoside (NADH) or its model compounds to substrates have been the subject of considerable controversy whether a hydride transfer occurs in a one-step[6-13] or consists of overall transfer of two electrons and a proton in a e^-–H^+–e^- sequence.[14-19] However, such mechanistic distinction

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seems to be actually dubious and possibly even meaningless unless either mechanism provides a more quantitative description of the energetic profiles of the reactions than the other does.

We have previously reported a quantitative approach to delineate the energetic profiles of hydride-transfer reactions from an NADH model compound, 1-benzyl-1,4-dihydronicotinamide (BNAH), to a series of *p*-benzoquinone derivatives (Q) in the absence and presence of Mg^{2+} ion in acetonitrile (MeCN) as proceeding via e^-–H^+–e^- sequence by showing that the rate constants and the primary kinetic isotope effects can be successfully correlated with the Gibbs energy change of the electron transfer from BNAH to Q.[20,21] Such a hydride transfer reaction via e^-–H^+–e^- sequence proceeds in the pair and should be distinguished from an electron-transfer reaction when the pair separates to yield free radical ion, $BNAH^{\bullet+}$ and $Q^{\bullet-}$.

Recently, Miller et al.[22] have reported an alternative elegant approach to describe the energetic profile of hydride-transfer reactions from NADH to Q in H_2O as proceeding by one-step hydride transfer by showing that the rate constants are related with the two-electron redox potentials for the Q/QH^- couple with a slope of 16.9 V^{-1} and ruled out an electron-transfer mechanism based on the comparison of data with the one-electron-transfer reactions from NADH to ferricenium ions (Fc^+) by assuming implicitly that the work terms w_p which are required to bring the product ions to the mean separation in the activated complex are the same between the NADH–Q and NADH–Fc^+ systems.[22,23] However, the work term w_p of the radical ion pair [$BNAH^{\bullet+}$ $Q^{\bullet-}$] in MeCN which consists of the opposite charges could be largely negative and thereby quite different from w_p of the radical pair [$NADH^{\bullet+}$ Fc] in H_2O which contains a neutral species Fc. The importance of the work term w_p in electron-transfer processes has been pointed out,[24] especially in the case of those via charge-transfer (CT) complexes formed between neutral donor and acceptor molecules.[2] In fact, the CT complexes formed between an NADH model compound and *p*-benzoquinone derivatives have been isolated, and the transient CT spectra have been detected during the hydride-transfer reactions from BNAH to Q.[20]

We report here the systematic kinetic studies on electron-transfer reactions from NADH model compounds to the excited states of $[RuL_3]^{2+}$ and $[RuL_2L']$ (L = 2,2′-bipyridine and L′ = 4,4′-$(COO^-)_2$bpy) which have different charges and those from the excited states of BNAH to various neutral organic oxidants by considering the difference of the work term w_p as well as on hydride-transfer reactions from various NADH model compounds to *p*-benzoquinone derivatives by determining independently the energetics of each step of the hydride-transfer reactions, i.e., the Gibbs energy changes of the e^-–H^+–e^- sequence. We wish to compare the two approaches to describe the energetic profiles of the hydride-transfer reactions; one is the correlation of the rate constants for both the photoinduced electron-transfer and hydride-transfer reactions with the Gibbs energy change of the electron transfer $\Delta G°_{et}$, and the other is the correlation of the rate constants of the hydride-transfer reactions with the Gibbs energy change of the overall transfer of a hydride ion $\Delta G°_{H^-}$ which is equal to the sum of the Gibbs energy changes of the e^-–H^+–e^- sequence. Different types of dihydropyridine (PyH_2) such as 1-(substituted benzyl)-1,4-dihydronicotinamide (X-BNAH) and *N*-methylacridan (AcH_2) which have similar one-electron redox potentials $E°(PyH_2^{\bullet+}/PyH_2)$ but different two-electron redox potentials $E°(PyH^+/PyH_2)$ are chosen as NADH model compounds so that it can be determined which factor ($\Delta G°_{et}$ or $\Delta G°_{H^-}$) reflects the change of the activation barrier of the hydride-transfer reactions better. *p*-Benzoquinone derivatives may be the best choice as the substrates for the purpose of this study since the thermodynamic data concerning the quinone–hydroquinone systems are well characterized.[25] We report also the role of Mg^{2+} ion on both the photoinduced electron-transfer and hydride-transfer reactions of NADH model compounds.

Experimental Section

Materials. Dihydropyridine derivatives, 1-benzyl-1,4-dihydronicotinamide (BNAH),[26] 1-(X-benzyl)-1,4-dihydronicotinamides (X-BNAH: X = 4-MeO, 4-Me, 4-Cl, and 2,4-Cl_2),[26,27] 2,6-dimethyl-3,5-dicarbethoxy-1,4-dihydropyridine (Hantzsch's ester),[28] 1-benzyl-3-carbamoyl-1,4-dihydroquinoline (BCQH),[29] and *N*-methylacridan (AcH_2)[11] were prepared according to the literatures. All the pyridinium cations (PyH^+) which are the oxidized forms of dihydropyridine derivatives were obtained as the perchlorate salts by the addition of magnesium perchlorate to the halide salts in water. The elemental analyses of all the dihydropyridine derivatives gave satisfactory results. Tris(2,2′-bipyridine)ruthenium(II) dichloride hexahydrate $[RuL_3]Cl_2·6H_2O$,[30] bis(2,2′-bipyridine)-4,4′-dicarbethoxy-2,2′-bipyridineruthenium(II) hexafluorophosphate, which dissociates to give $[RuL_2L']$ (L′ = 4,4′-$(COO^-)_2$bpy) in MeCN,[31] and tris(2,2′-bipyridine)iron(III) perchlorate $[FeL_3](ClO_4)_3$[2a] were prepared according to the literatures. Organic oxidants (*N*,*N*-dimethylaniline, *N*,*N*-dimethyl-*p*-toluidine, *N*,*N*′-diphenyl-*p*-phenylenediamine) and oxidants (1-cyanonaphthalene, *trans*-stilbene, and cyanobenzene) used as quenchers for the luminescence quenching of $[RuL_3]^{2+}$ and $[RuL_2L']$ were obtained commercially and purified by the standard method.[32] Most *p*-benzoquinone derivatives (2,3-dichloro-5,6-dicyano-*p*-benzoquinone, *p*-chloranil, *p*-bromanil, 2,6-dichloro-*p*-benzoquinone, *p*-benzoquinone, methyl-*p*-benzoquinone, and 2,6-dimethyl-*p*-benzoquinone) were also obtained commercially and purified by the standard methods.[32] Chloro-*p*-benzoquinone and 2,3-dicyano-*p*-benzoquinone were prepared from the corresponding hydroquinones according to the literature.[33] Anhydrous magnesium perchlorate was obtained from Nakarai Chemicals. Acetonitrile which was also obtained commercially was purified and dried with calcium hydride by the standard procedure[32] and stored under nitrogen atmosphere.

Luminescence Quenching. Quenching experiments of the BNAH fluorescence and the $[RuL_3]^{2+}$ and $[RuL_2L']$ luminescence were performed by using a Hitachi 650-10S fluorescence spectrophotometer. The excitation wavelengths were 350, 363, 452, and 470 nm for BNAH in the absence of Mg^{2+} ion, BNAH in the presence of Mg^{2+} ion, and $[RuL_3]^{2+}$ and $[RuL_2L']$ in MeCN, respectively. The monitoring wave-

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lengths were those corresponding to the maxima of the respective emmision bands at 443, 455, 608, and 630 nm. Relative emission intensities were measured for MeCN solutions of BNAH, $[RuL_3]^{2+}$, and $[RuL_2L']$ containing n-Bu_4NClO_4 (1.0×10^{-2} or 0.10 M) with a quencher at various concentrations. The solutions of $[RuL_3]^{2+}$ and $[RuL_2L']$ were degassed and replaced by an atmospheric pressure of nitrogen prior to the measurements. There has been no change in the shape except intensity of the emission spectrum by the addition of a quencher. The Stern–Volmer relationship (eq 1) was obtained between the ratio of the

$$I_f^0/I_f = 1 + K_q[q] \quad (1)$$

emission intensities in the absence and presence of a quencher I_f^0/I_f and the quencher concentration $[q]$. The observed rate constant k_{obsd} ($= K_q\tau^{-1}$) was determined from the Stern–Volmer constants K_q and the emission lifetime τ (BNAH 0.76 ns,[34] BNAH in the presence of Mg^{2+} ion 1.03 ns,[35] $[RuL_3]^{2+}$ 850 ns,[36] and $[RuL_2L']$ 850 ns[36]).

Cyclic Voltammetry. Reduction potentials of $[RuL_2L']$ and organic oxidants which have not been reported were determined by the cyclic voltammetry measurements, which were performed on a Hokuto Denko Model HA-301 potentiostat-galvanostat at 298 K in MeCN containing 0.10 M n-Bu_4NClO_4 as a supporting electrolyte by using a saturated calomel electrode (SCE) or an Ag/Ag^+ electrode ($Ag/AgNO_3$ 0.10 M) as a reference one under deaerated conditions; the redox potential vs. Ag/Ag^+ is more negative by 0.34 V than that vs. SCE.[37] The one-electron redox potentials of PyH_2 $E°(PyH_2^{\bullet+}/PyH_2)$ and PyH^+ $E°(PyH^+/PyH^\bullet)$ in the absence and presence of Mg^{2+} ion were determined by analyzing the cyclic voltammograms at various sweep rates in the range 10–1000 mV s^{-1}, based on the method described in detail elsewhere.[21,38] The platinum microelectrode was routinely cleaned by soaking it in concentrated nitric acid, followed by repeating rinsing with water and acetone, and drying at 353 K prior to use. The anodic and cathodic peak potentials of PyH_2 E_{ox}^p and PyH^+ E_{red}^p, respectively, were reproducible to within ±20 mV at a constant sweep rate. Little deterioration of the electrode was observed upon repeated scans, indicating that anodic oxidation of PyH_2 or cathodic reduction of PyH^+ leads to products without fouling of the electrode surface.

Kinetic Measurements. Kinetic measurements were performed under deaerated conditions by using a Union RA-103 stopped flow spectrophotometer and a Union SM-401 spectrophotometer for fast reactions of PyH_2 with $[Fe(bpy)_3]^{3+}$ and p-benzoquinone derivatives (Q) with half-lives shorter than 10 s and for the slower reactions of PyH_2 with Q with half-lives much longer than 10 s, respectively. Rates of the electron-transfer reactions from PyH_2 to $[Fe(bpy)_3]^{3+}$ in MeCN containing an excess amount of pyridine derivative to PyH_2 were followed by the increase in absorbance at λ_{max} = 507 nm due to $[Fe(bpy)_3]^{2+}$.[39] Rates of the hydride-transfer reactions from X-BNAH to Q in the absence and presence of Mg^{2+} ion in MeCN at 298 K were determined by the increase in absorbance at the absorption maxima of the respective semiquinone radical anions[20] and by the decrease in absorbance at λ_{max} of X-BNAH, respectively. Rates of the hydride-transfer reactions from AcH_2 to Q in both the absence and presence of Mg^{2+} ion in MeCN at 298 K were monitored by the increase in absorbance due to AcH^+ in the visible region (λ_{max} 358 nm, ϵ 2.00×10^4 M^{-1} cm^{-1}). All the kinetic measurements were carried out under the pseudo-first-order conditions by using more than 10-fold excess substrates. Especially, in the case of the hydride-transfer reaction from AcH_2 to chloro-p-benzoquinone in the absence of Mg^{2+} ion in MeCN at 298 K, more than 500-fold excess of the quinone was used in order to neglect the back reaction. The pseudo-first-order rate constants were determined by the least-square curve fit, by using a Union System 77 microcomputer.

Results

Photoinduced Electron-Transfer Reactions of NADH Model Compounds. The luminescence of $[RuL_3]^{2+*}$ (λ_{max} 608 nm, τ 850 ns) is known to be quenched by the electron-transfer reactions from various electron donors to $[RuL_3]^{2+*}$.[36,40] When the absorption band of $[RuL_3]^{2+}$ in MeCN containing BNAH was excited at λ_{max} 452 nm where BNAH has no absorption band, BNAH efficiently quenched the $[RuL_3]^{2+*}$ luminescence by the electron transfer from BNAH to $[RuL_3]^{2+*}$ (eq 2). The observed

$$[RuL_3]^{2+*} + BNAH \xrightarrow{k_{obsd}} [RuL_3]^+ + BNAH^{\bullet+} \quad (2)$$

rate constant k_{obsd} in a deaerated MeCN solution containing 1.0×10^{-2} M n-Bu_4NClO_4 is determined as 4.1×10^8 M^{-1} s^{-1} from the Stern–Volmer plot in Figure 1a. The k_{obsd} values for various NADH model compounds are listed in Table I. NADH model compounds quenched also the luminescence of $[RuL_2L']^*$ (λ_{max} 630 nm, τ 850 ns)[36] in MeCN containing 1.0×10^{-2} M n-Bu_4NClO_4 by the electron-transfer reactions from NADH model compounds to $[RuL_2L']^*$ (eq 3). The k_{obsd} values are also listed

$$[RuL_2L']^* + BNAH \xrightarrow{k_{obsd}} [RuL_2L']^- + BNAH^{\bullet+} \quad (3)$$

in Table I, where the k_{obsd} values in both cases (eq 2 and 3) decrease with decreasing the donor ability of the substituent on BNAH.

When $Mg(ClO_4)_2$ was added to the $[RuL_3]^{2+*}$–BNAH system, k_{obsd} value decreased remarkably with an increase of the Mg^{2+} concentration as shown in Figure 1a. A similar retarding effect of Mg^{2+} ion was observed in the reductive quenching of $[RuL_3]^{2+*}$ by other NADH model compounds; the k_{obsd} values in the presence

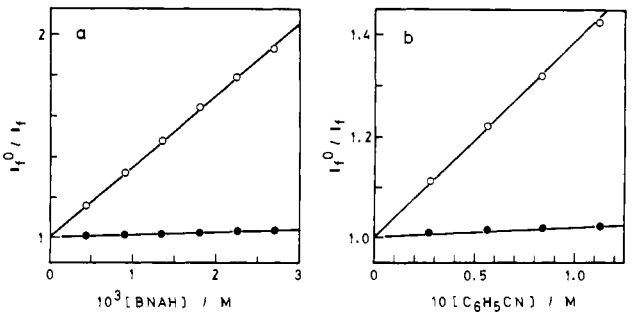


Figure 1. Stern–Volmer plots (a) for the luminescence quenching of $[RuL_3]^{2+*}$ (1.3×10^{-5} M) by BNAH in the absence (○) and presence of 1.0×10^{-2} M $Mg(ClO_4)_2$ (●) in MeCN containing 1.0×10^{-2} M n-Bu_4NClO_4 and (b) for the fluorescence quenching of BNAH* (1.8×10^{-4} M) by cyanobenzene in the absence (○) and presence of 5.0×10^{-2} M $Mg(ClO_4)_2$ (●) in MeCN at 298 K.

Table I. Rate Constants k_{obsd} for the Luminescence Quenching of $[RuL_3]^{2+*}$ and $[RuL_2L']^*$ by Organic Reductants (PyH_2 and Aromatic Amines) in MeCN at 298 K

	k_{obsd} (M^{-1} s^{-1})[a] in the presence of 1.0×10^{-2} M			
reductant	$[RuL_3]^{2+*}$ n-Bu_4NClO_4	$[RuL_3]^{2+*}$ $NaClO_4$	$[RuL_3]^{2+*}$ $Mg(ClO_4)_2$	$[RuL_2L']^*$ n-Bu_4NClO_4
4-MeOBNAH	8.5×10^8	7.3×10^8	8.6×10^7	2.6×10^9
4-MeBNAH	5.5×10^8	5.5×10^8	8.6×10^7	1.7×10^9
BNAH	4.1×10^8	3.8×10^8	1.5×10^7	1.1×10^9
4-ClBNAH	2.8×10^8	2.7×10^8	1.2×10^7	6.8×10^8
2,4-Cl_2BNAH	2.1×10^8	1.7×10^8	1.2×10^7	6.0×10^8
Hantzsch's ester	1.2×10^8	1.1×10^8	$<10^6$	
BCQH	6.3×10^7		$<10^6$	
N,N-dimethyl-aniline	9.4×10^7[b]	8.5×10^7	9.1×10^7	
N,N-dimethyl-p-toluidine	9.4×10^8[b]	8.7×10^8	8.8×10^8	
N,N'-diphenyl-p-phenylene-diamine	5.6×10^9[b]	5.7×10^9	5.4×10^9	

[a] The experimental errors are within ±5%. [b] Similar values are reported in ref 40.

(34) Martens, F. M.; Verhoeven, J. W.; Gase, R. A.; Pandit, U. K.; de Bore, Th. J. *Tetrahedron* **1978**, *34*, 443.
(35) The fluorescence lifetime τ of BNAH* in the presence of Mg^{2+} ion is estimated from the K_q value for diethyl fumarate, which is assumed to be diffusion limited (k_{diff} = 2.0×10^{10} M^{-1} s^{-1}).
(36) (a) Boch, C. R.; Conner, J. A.; Gutierrez, A. R.; Meyer, T. J.; Whitten, D. G.; Sullivan, B. P.; Nagle, J. K. *J. Am. Chem. Soc.* **1979**, *101*, 4815. (b) The luminescence lifetime τ of $[RuL_2L']^*$ in MeCN is taken to be the same as that of $[RuL_3]^{2+*}$ in MeCN since their τ values in D_2O are identical (ref 31b).
(37) Mann, C. K. *Electroanal. Chem.* **1969**, *4*, 57.
(38) Fukuzumi, S.; Hironaka, K.; Nishizawa, N.; Tanaka, T. *Bull. Chem. Soc. Jpn.* **1983**, *56*, 2220.
(39) Fukuzumi, S.; Nishizawa, N.; Tanaka, T. *Bull. Chem. Soc. Jpn.* **1982**, *55*, 3482.
(40) Ballardini, R.; Varani, G.; Indelli, M. T.; Scandola, F.; Balzani, V. *J. Am. Chem. Soc.* **1978**, *100*, 7219.

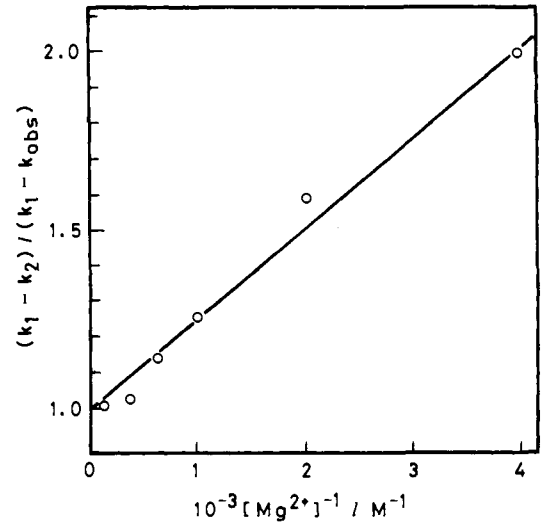


Figure 2. Plot of $(k_1 - k_2)/(k_1 - k_{obsd})$ vs. $[Mg^{2+}]^{-1}$ for the luminescence quenching of $[RuL_3]^{2+*}$ by the electron transfer from BNAH to $[RuL_3]^{2+*}$ in the presence of Mg^{2+} ion in MeCN at 298 K.

of 1.0×10^{-2} M Mg^{2+} ion are shown in Table I. Such a retarding effect may be caused by the specific interaction of NADH model compounds with Mg^{2+} ion, since no effect of Mg^{2+} ion has been observed on the $[RuL_3]^{2+*}$ luminescence quenching by aromatic amines (*N*,*N*-dimethylaniline, *N*,*N*-dimethyl-*p*-toluidine, and *N*,*N*'-diphenyl-*p*-phenylenediamine) as quenchers and no retarding effect has been detected in the presence of other metal ion (1.0×10^{-2} M $NaClO_4$) either (Table I). In fact, BNAH is known to form a 1:1 complex with Mg^{2+} ion (eq 4) with the formation

$$\text{BNAH} + Mg^{2+} \overset{K_1}{\rightleftharpoons} [\text{BNAH } Mg^{2+}] \qquad (4)$$

constant $K_1 = 1.2 \times 10^4$ M^{-1} in MeCN at 298 K.[41] In the presence of Mg^{2+} ion, the electron transfer from both BNAH and $[\text{BNAH } Mg^{2+}]$ to $[RuL_3]^{2+*}$ may occur, and then the observed rate constant k_{obsd} in the presence of Mg^{2+} ion may be expressed by eq 5 where k_1 and k_2 are the rate constants in the absence of

$$k_{obsd} = \frac{k_1 + k_2K_1[Mg^{2+}]}{1 + K_1[Mg^{2+}]} \qquad (5)$$

Mg^{2+} ion and in the presence of a large excess Mg^{2+} ion such that all BNAH molecules form the 1:1 complex with Mg^{2+} ion, respectively. Equation 5 is rewritten by eq 6, the validity of which

$$\frac{(k_1 - k_2)}{(k_1 - k_{obsd})} = 1 + \frac{1}{K_1[Mg^{2+}]} \qquad (6)$$

is shown by the linear plot of $(k_1 - k_2)/(k_1 - k_{obsd})$ vs. $[Mg^{2+}]^{-1}$ in Figure 2. The K_1 value which is obtained from the reciprical of the slope in Figure 2 ($K_1 = 3.6 \times 10^3$ M^{-1}) is somehow smaller than the value which was reported previously,[41] since the presence of water contained in $[RuL_3]Cl_2 \cdot 6H_2O$ is known to decrease the K_1 value in MeCN.[41] Thus, the observed rate constant k_{obsd} in the presence of large excess Mg^{2+} ion corresponds to the rate constants for the electron transfer from $[\text{BNAH } Mg^{2+}]$ to $[RuL_3]^{2+*}$.

The excitation of the absorption band (λ_{max} 350 nm) of an MeCN solution of BNAH results in a fluorescence (λ_{max} 443 nm, τ 0.76 ns).[34] The BNAH fluorescence is known to be quenched by the electron transfer from BNAH* to organic oxidants (ox) in MeCN (eq 7).[34,39] A retarding effect of Mg^{2+} ion is observed

$$\text{BNAH}^* + \text{ox} \xrightarrow{k_{obsd}} \text{BNAH}^{\bullet+} + \text{ox}^{\bullet-} \qquad (7)$$

also in the oxidative quenching of BNAH* by an organic oxidant in MeCN as shown in Figure 1b. The k_{obsd} values for various organic oxidants in the absence and presence of 5.0×10^{-2} M Mg^{2+} ion are listed in Table II, together with the one-electron reduction potentials of organic oxidants. The k_{obsd} value in the absence of Mg^{2+} ion increases with the positive shift of the one-electron reduction potentials of organic oxidants to reach the value close to diffusion limited. In the presence of 5.0×10^{-2} Mg^{2+} ion, the k_{obsd} values are smaller than those in its absence, except for the case of diethyl fumarate in which the rate constants in both the absence and presence of Mg^{2+} ion are close to diffusion limited as shown in Table II, where the retarding effect of Mg^{2+} ion is pronounced in the case of organic oxidants (1-cyanonaphthalene, *trans*-stilbene, and cyanobenzene) which have no carbonyl group to interact with Mg^{2+} ion.

Table II. Rate Constants k_{obsd} for the Fluorescence Quenching of BNAH* by Organic Oxidants in the Absence and Presence of Mg^{2+} Ion and Their One-Electron Reduction Potentials $E°(ox/ox^{\bullet-})$ (vs. SCE) in MeCN

oxidant	$E°(ox/ox^{\bullet-})$ V	k_{obsd} (M^{-1} s^{-1})[c] at the Mg^{2+} concentration (M)	
		0[d]	5.0×10^{-2}
diethyl fumarate	−1.50[a]	1.4×10^{10}	2.0×10^{10}
diethyl terephthalate	−1.78[a]	1.7×10^{10}	
benzophenone	−1.86[a]	1.7×10^{10}	
1-cyanonaphthalene	−1.96[a]	1.6×10^{10}	6.1×10^{9}
acetophenone	−2.10[a]	9.1×10^{9}	8.1×10^{9}
propiophenone	−2.15[b]	9.1×10^{9}	
trans-stilbene	−2.20[a]	5.8×10^{9}	8.1×10^{8}
4-methylacetophenone	−2.20[b]	6.7×10^{9}	4.9×10^{9}
4-methoxyacetophenone	−2.22[b]	5.9×10^{9}	4.1×10^{9}
cyanobenzene	−2.35[a]	4.9×10^{9}	2.5×10^{8}
methyl benzoate	−2.37[a]	4.4×10^{9}	7.4×10^{8}
ethyl benzoate	−2.40[a]	3.6×10^{9}	1.4×10^{9}

[a] Taken from ref 34. [b] Determined by the cyclic voltammograms (see Experimental Section). [c] The experimental errors are within ±5%. [d] Taken from ref 38.

Redox Potentials of PyH_2 and PyH^+. The one-electron oxidation and reduction potentials of PyH_2 $E°(PyH_2^{\bullet+}/PyH_2)$ and PyH^+ $E°(PyH^+/PyH^{\bullet})$, respectively, in the absence and presence of Mg^{2+} ion in MeCN were determined by analyzing the cyclic voltammograms at various sweep rates, based on the method to determine the redox potentials in irreversible systems,[21,38] as follows. The cyclic voltammograms of PyH_2 in the absence and presence of of Mg^{2+} ion are characterized by well-defined anodic waves, but no cathodic waves were observed on the reverse scan, even at high scan rates. Conversely, those of PyH^+ show well-defined cathodic waves but no corresponding anodic waves. The irreversible behavior suggests that the radical species formed on anodic oxidation or cathodic reduction is unstable and the follow-up chemical reaction is fast on the time scale of the CV measurements. Assuming that the follow-up chemical reaction is not rate-limiting, the width of the wave $E^p - E^{p/2}$ in such an irreversible system has been shown to depend only on the transfer coefficient β according to eq 8[1f,38,42] where F is the Faraday

$$E^p - E^{p/2} = 1.857(RT/\beta F) \qquad (8)$$

constant. It should be noted that not only E^p but also $E^p - E^{p/2}$ varies with the sweep rate.[42] Thus, the transfer coefficient β defined by the tangent of the Gibbs energy change of electron transfer $\partial\Delta G_{et}^{\ddagger}/\partial\Delta G°_{et}$ at the peak potential is obtained from the width of the wave $E^p - E^{p/2}$ by using eq 9, which is derived from eq 8. On the other hand, the activation Gibbs energy of electron

$$\beta = \frac{1.857RT}{F(E^p - E^{p/2})} \qquad (9)$$

transfer $\Delta G_{et}^{\ddagger}$ is expressed as a function of the standard Gibbs

(41) Fukuzumi, S.; Kondo, Y.; Tanaka, T. *Chem. Lett.* **1983**, 485.

(42) (a) Nicholson, R. S.; Shain, I. *Anal. Chem.* **1964**, *36*, 706. (b) Klingler, R. J.; Kochi, J. K. *J. Am. Chem. Soc.* **1980**, *102*, 4790. (c) Tamblyn, W. H.; Klingler, R. J.; Hwang, W. S.; Kochi, J. K. *Ibid.* **1981**, *103*, 3161. (d) Klingler, R. J.; Kochi, J. K. *Ibid.* **1982**, *104*, 4186. (e) Klingler, R. J.; Kochi, J. K. *Ibid.* **1981**, *103*, 5839. (f) Klingler, R. J.; Kochi, J. K. *J. Phys. Chem.* **1981**, *85*, 1731.

Table III. One-Electron Oxidation and Reduction Potentials (vs. SCE) of PyH_2 and PyH^+, Respectively, in the Absence and Presence of Mg^{2+} Ion in MeCN at 298 K[a]

PyH_2	$E°(PyH_2^{\cdot+}/PyH_2)$[b] V	$E°(PyH_2^{\cdot+}/PyH_2)$[c] V
4-MeOBNAH	0.50	0.68
4-MeBNAH	0.54	0.70
BNAH	0.57	0.80
4-ClBNAH	0.62	0.77
2,4-Cl_2BNAH	0.59	0.78
Hantzsch's ester	0.72	0.94
AcH_2	0.80	0.79

PyH^+	$E°(PyH^+/PyH^{\cdot})$[b] V
4-MeBNA$^+$	−1.13
BNA$^+$	−1.08
4-ClBNA$^+$	−1.08
2,4-Cl_2BNA$^+$	−1.08
AcH^+	−0.43

[a] Determined by the analysis of the cyclic voltammograms (see text and ref 21 and 38). The experimental errors are within ±5%. [b] In the presence of 0.10 M n-Bu_4NClO_4. [c] In the presence of 5.0×10^{-2} M $Mg(ClO_4)_2$ and n-Bu_4NClO_4.

energy change of electron transfer $\Delta G°_{et}$ by using Marcus equation (eq 10)[43] equal to the activation Gibbs energy when $\Delta G°_{et} = 0$.

$$\Delta G_{et}^{\ddagger} = \Delta G_0^{\ddagger}\left(1 + \frac{\Delta G°_{et}}{4\Delta G_0^{\ddagger}}\right)^2 \quad (10)$$

Then, the transfer coefficient β is given as the function of $\Delta G°_{et}$ by differentiating eq 10 with respect to $\Delta G°_{et}$ (eq 11). From eq

$$\beta = \frac{1}{2} + \frac{1}{8}\frac{\Delta G°_{et}}{\Delta G_0^{\ddagger}} \quad (11)$$

11 are derived relations of oxidation and reduction peak potentials with the standard oxidation and reduction potentials when $\Delta G°_{et} = E°(PyH_2^{\cdot+}/PyH_2) - E^p(PyH_2^{\cdot+}/PyH_2)$ (eq 12) and $\Delta G°_{et} = E^p(PyH^+/PyH^{\cdot}) - E°(PyH^+/PyH^{\cdot})$ (eq 13), respectively.

$$E^p(PyH_2^{\cdot+}/PyH_2) = E°(PyH_2^{\cdot+}/PyH_2) + 4(1 - 2\beta)\Delta G_0^{\ddagger} \quad (12)$$

$$E^p(PyH^+/PyH^{\cdot}) = E°(PyH^+/PyH^{\cdot}) - 4(1 - 2\beta)\Delta G_0^{\ddagger} \quad (13)$$

The $E^p(PyH_2^{\cdot+}/PyH_2)$ and $E^p(PyH^+/PyH^{\cdot})$ values at various sweep rates are plotted against $4(1 - 2\beta)$ in which the β values are obtained from the width of the wave $E^p - E^{p/2}$ by using eq 9. Typical examples of the plots $E^p(PyH_2^{\cdot+}/PyH_2)$ and $E^p(PyH^+/PyH^{\cdot})$ vs. $4(1 - 2\beta)$ are shown in Figure 3 (parts a and b, respectively) where the observed linear correlations agree well with eq 12 and 13. Thus, from the intercepts of the linear plots, the $E°(PyH_2^{\cdot+}/PyH_2)$ and $E°(PyH^+/PyH^{\cdot})$ values are determined as listed in Table III, which shows that the one-electron oxidation potential of X-BNAH is shifted to the positive direction with decreasing the donor ability of the substituent X. The one-electron oxidation potentials of various NADH model compounds in MeCN (0.50–0.80 V) are similar to that of NADH in H_2O reported by Miller et al.[44] (0.69 V),[45] while the one-electron reduction potentials of X-BNA$^+$ (−1.1 V) are more negative than that of NAD$^+$ (−0.70 V).[45,46] An important point to note from Table III is that the one-electron reduction potential of AcH^+ (−0.43 V) is much more positive than those of X-BNA$^+$ (−1.1 V) while the one-electron oxidation potential of AcH_2 (0.80 V) is only slightly more positive than those of X-BNAH (0.50–0.67 V). Such a positive shift of the reduction potential of AcH^+ relative to X-BNA$^+$ may be ascribed to the delocalization of the unpaired electron of $AcH^{\cdot}$. A similar one-electron reduction potential has been reported in the case of *N*-methyl-9-phenyl-acridinium ion (−0.46 V) which shows a reversible cyclic voltammogram.[47] It should also be noted that the one-electron oxidation potential of X-BNAH in the presence of Mg^{2+} (5.0×10^{-2} M) in MeCN is shifted to the positive direction (ca. +0.2 V) relative to that in its absence by forming the complex with Mg^{2+} ion (eq 4) and shows no further shift in the presence of much more concentrated Mg^{2+} ion (1.6 M),[21] while that of AcH_2 which has no carbonyl group to interact with Mg^{2+} ion shows essentially no shift in the presence of Mg^{2+} ion (Table III).

Scheme I

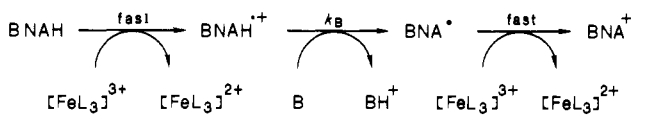


Table IV. Rate Constants k_B for the Deprotonation of the Radical Cations of PyH_2 by Pyridine Derivatives in MeCN at 298 K and the pK_a Values of Pyridine Derivatives

		log k_B (M^{-1} s^{-1})			
pyridine derivative	pK_a[a]	BNAH[a]	4-ClBNAH	2,4-Cl_2BNAH	AcH_2
3,5-dichloropyridine	0.67	3.2	4.0	3.8	4.9
3-cyanopyridine	1.45	3.8	4.4	4.4	5.4
4-cyanopyridine	1.86	4.1	5.0	4.8	6.0
3-bromopyridine	2.84	5.0	5.7	5.5	6.2
3-acetylpyridine	3.18	5.4	5.9	5.7	6.1
4-acetylpyridine	3.51	5.6	6.1	5.8	6.2
pyridine	5.29	6.0	6.2	5.9	6.3
2-aminopyridine	6.82	6.1			

[a] Taken from ref 17a.

Table V. pK_a Values of Radical Cations of NADH Model Compounds

NADH model compd	pK_a
BNAH[a]	3.6
4-ClBNAH	3.4
2,4-Cl_2BNAH	3.3
AcH_2	2.0

[a] Taken from ref 17a.

pK_a Values of Radical Cations of PyH_2. The pK_a value of $BNAH^{\cdot+}$ has previously been determined as 3.6 from a Brønsted plot of the logarithm of the rate constant k_B for the proton transfer from $BNAH^{\cdot+}$ to various bases vs. pK_a of the bases.[17] The proton-transfer rate constant k_B can be determined by analyzing the kinetics of the electron-transfer reactions from BNAH to $[FeL_3]^{3+}$ (L = 2,2′-bipyridine) in MeCN containing an excess amount of base as follows. In the presence of a base, BNAH acts as an apparent two-electron donor, when the two-electron transfer proceeds via a two-step process; first, a fast one-electron transfer from BNAH to $[FeL_3]^{3+}$ occurs,[48] followed by the rate-determining deprotonation of $BNAH^{\cdot+}$ by a base (B), and the subsequent second electron transfer from BNA$^{\cdot}$ to $[FeL_3]^{3+}$ is also very fast[49] (Scheme I). Thus, the rate of formation of $[FeL_3]^{2+}$ for the second step of the electron-transfer reaction from BNAH

(43) Marcus, R. A. *J. Phys. Chem.* **1963**, *67*, 853. Marcus, R. A. *Ann. Rev. Phys. Chem.* **1964**, *15*, 155.
(44) Carlson, B. W.; Miller, L. L.; Neta, P.; Grodkowski, J. *J. Am. Chem. Soc.* **1984**, *106*, 7233.
(45) The reported potential vs. NHE is converted to that vs. SCE by subtracting 0.24 V.
(46) (a) Farrington, J. A.; Land, E. J.; Swallow, A. J. *Biochim. Biophys. Acta* **1980**, *590*, 273. (b) Anderson, R. *Ibid.* **1980**, *590*, 277.
(47) Koper, N. W.; Verhoeven, J. W.; Jonker, S. A. In *Proceedings of Xth IUPAC Symposium on Photochemistry*; Presses Polytechniques Romandes: Lausanne, 1984, p 323.
(48) The Gibbs energy change of the electron transfer from BNAH to $[FeL_3]^{3+}$ is evaluated as −11 kcal mol^{-1} from the one-electron oxidation potential of the reductant BNAH (0.57 V in Table III) and the one-electron reduction potential of $[FeL_3]^{3+}$ (1.06 V)[39] by using eq 33.
(49) The Gibbs energy change of the electron transfer from BNA$^{\cdot}$ to $[FeL_3]^{3+}$ is evaluated as −49 kcal mol^{-1} from the one-electron oxidation potential of the reductant BNA$^{\cdot}$, which is equivalent to the one-electron reduction potential of BNA^+ (−1.08 V in Table III) and the one-electron reduction potential of $[FeL_3]^{3+}$ (1.06 V)[39] by using eq 33.

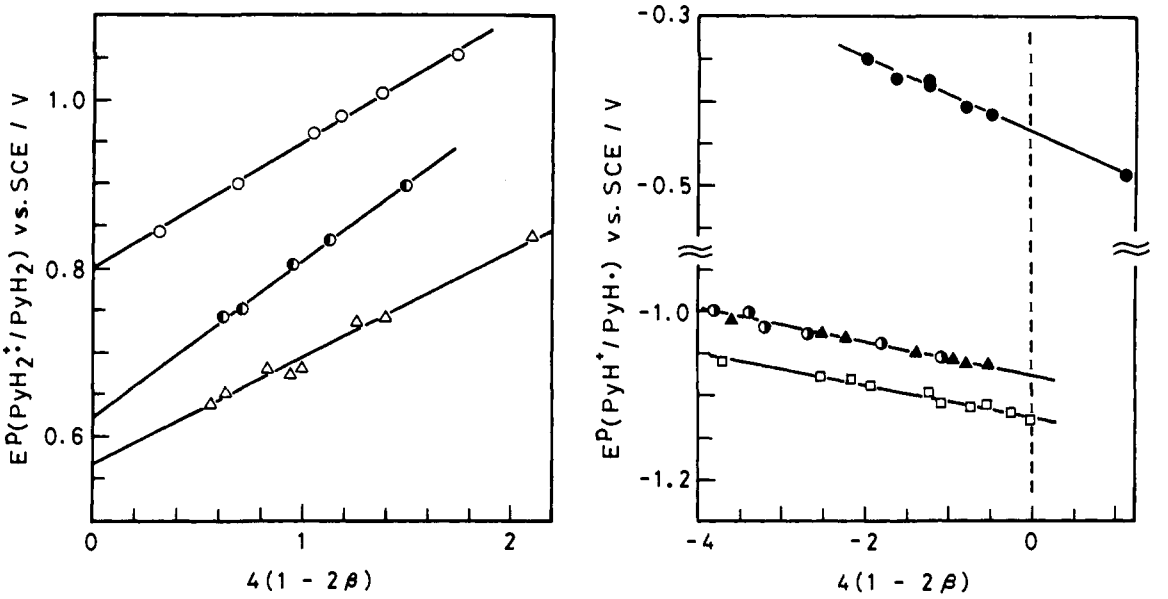


Figure 3. (a) Oxidation peak potentials of PyH_2, $E^p(PyH_2^{\bullet+}/PyH_2)$, and (b) reduction peak potentials of PyH^+, $E^p(PyH^+/PyH^\bullet)$ in MeCN at 298 K plotted as a function of transfer coefficient β, $4(1 - 2\beta)$, see eq 12 and 13, respectively: AcH_2 (○), 4-ClBNAH (◐), BNAH (△), AcH^+ (●), $2,4\text{-}Cl_2BNA^+$ (◑), BNA^+ (▲), and $4\text{-}MeBNA^+$ (□).

Table VI. Rate Constants k_{obsd} for the Hydride-Transfer Reactions from NADH Model Compounds to *p*-Benzoquinone Derivatives (Q) and the One-Electron Reduction Potentials of Q in the Absence and Presence of Mg^{2+} Ion in MeCN at 298 K

			log k_{obsd}[a] (M^{-1} s^{-1})											
	$E°(Q/Q^{\bullet-})$ V vs. SCE $[Mg^{2+}]$/M		BNAH[b] $[Mg^{2+}]$/M		4-MeOBNAH $[Mg^{2+}]$/M		4-MeBNAH $[Mg^{2+}]$/M		4-ClBNAH $[Mg^{2+}]$/M		$2,4\text{-}Cl_2BNAH$ $[Mg^{2+}]$/M		AcH_2 $[Mg^{2+}]$/M	
p-benzoquinone derivative	0	0.10	0	0.10	0	0.10	0	0.10	0	0.10	0	0.10	0	0.10
2,3-dichloro-5,6-dicyano-*p*-benzoquinone	0.51	0.51	6.92										6.18	6.15
2,3-dicyano-*p*-benzoquinone	0.28	0.25	5.86	4.45	6.23	4.52					5.48	3.83	4.04	4.28
p-chloranil	0.01	0.03	3.00	1.26	3.28	1.36	3.20				2.62	1.86	1.08	1.23
p-bromanil	0	−0.04	2.86	1.38	3.11	1.51	3.96	1.40	2.70		2.30	1.00	0.95	1.18
2,6-dichloro-*p*-benzoquinone	−0.18	−0.04	1.88	0.71	2.08	1.04	1.99	0.76	1.65	0.51	1.63	0.30	−0.82	−0.09
2,5-dichloro-*p*-benzoquinone	−0.18	−0.04	1.70	0.48	1.77	0.65	1.64	0.49	1.41	0.28	1.15	0.08		
chloro-*p*-benzoquinone	−0.34	−0.13	0.88	0	0.69	0.15	0.76	0.08	0.36	−0.21	0.36	−0.47	−2.11	−0.92
p-benzoquinone	−0.50	−0.18	−1.89	−0.57										−1.80
methyl-*p*-benzoquinone	−0.58	−0.32	−2.64	−1.36										−2.36
2,6-dimethyl-*p*-benzoquinone	−0.67	−0.30	−4.08	−2.41										−3.01
trimethyl-*p*-benzoquinone	−0.75	−0.40	−4.89											−3.58

[a] The experimental errors of k_{obsd} are within ±5%. [b] Taken from ref 21.

to $[FeL_3]^{3+}$ in the presence of an excess base after the completion of the initial electron transfer is given by eq 14 where $[Fe^{2+}]_\infty$

$$d[Fe^{2+}]/dt = k_B[B]([Fe^{2+}]_\infty - [Fe^{2+}]) \quad (14)$$

is the final concentration of $[FeL_3]^{2+}$ which is equal to twice of the initial concentration of BNAH. The k_B values of the proton transfer from $PyH_2^{\bullet+}$ to a series of bases are listed in Table IV, together with the pK_a values of bases. Since proton transfers between normal acids and bases are known to give biphasic Brønsted plots with breaks at $\Delta pK_a = 0$,[50] the pK_a values of $PyH_2^{\bullet+}$ can be determined from the breaks in Brønsted plots of log k_B vs. pK_a as shown in Figure 4.[51] In the case of BNAH, it has been confirmed that the break of the Brønsted plot occurs at $\Delta pK_a \cong 0$, where the primary kinetic isotope effects show the maximum.[17a] The pK_a values of X-$BNAH^{\bullet+}$ and $AcH_2^{\bullet+}$ are listed in Table V, which shows that the pK_a value of AcH_2^+ (2.0) is less than those of X-$BNAH^{\bullet+}$ (3.3–3.6).[52]

Kinetics of Hydride-Transfer Reactions from PyH_2 to Q. It has previously been reported that the hydride transfer from NADH model compounds to *p*-benzoquinone derivatives Q (eq 15) occurs in MeCN, followed by a subsequent fast reaction of QH^- with Q to form $Q^{\bullet-}$ and QH_2 (eq 16).[20,53] The rate of formation of $Q^{\bullet-}$ in the presence of a large excess Q is given by eq 17,[20] where $[Q^{\bullet-}]_\infty$ is the final concentration of $Q^{\bullet-}$ which is

$$PyH_2 + Q \xrightarrow{k_{obsd}} PyH^+ + QH^- \quad (15)$$

$$2QH^- + Q \xrightarrow{fast} 2Q^{\bullet-} + QH_2 \quad (16)$$

$$d[Q^{\bullet-}]/dt = k_{obsd}[Q]([Q^{\bullet-}]_\infty - [Q^{\bullet-}]) \quad (17)$$

(50) (a) Westheimer, F. H. *Chem. Rev.* **1961**, *61*, 265. (b) Bell, R. P. *The Protons in Chemistry*; Cornell: Ithaca, New York, 1973. (c) More O'Ferrall, R. A. In *Proton Transfer Reactions*; Caldin, E. F., Gold, V., Eds.; Chapman and Hall: London, 1975; Chapter 8. (d) Cohen, A. O.; Marcus, R. A. *J. Phys. Chem.* **1968**, *72*, 4249. (e) Bergman, N.-A.; Chiang, Y.; Kresge, A. J. *J. Am. Chem. Soc.* **1978**, *100*, 5954.

(51) The plateau values in Figure 4 are significantly lower than the diffusion-limited value, suggesting that there is a work term required to form the encounter complex preceding the proton transfer as is usual in a number of proton-transfer processes (see: Kresge, A. J. *Acc. Chem. Res.* **1975**, *8*, 354).

(52) The pK_a value of $NADH^{\bullet+}$ in H_2O has been reported to be −3.5, which is significantly less than the pK_a values of $PyH_2^{\bullet+}$ in Table V, based on the thermodynamic calculation (ref 44; see, also: Martens, F. M.; Verhoeven, J. W. *Recl. Trav. Chim. Pays-Bas* **1981**, *100*, 228). However, the larger pK_a value of $NADH^+$ (≅1) has recently been reported from the direct detection of $NADH^{\bullet+}$ by means of nanosecond laser photolysis (see: Czochralska, B.; Lindqvist, L. *Chem. Phys. Lett.* **1983**, *101*, 297).

(53) (a) Colter, A. K.; Saito, G.; Sharom, F. J. *Can. J. Chem.* **1977**, *55*, 2741. (b) Colter, A. K.; Saito, G.; Sharom, F. J.; Hong, A. P. *J. Am. Chem. Soc.* **1976**, *98*, 7833. (c) Saito, G.; Colter, A. K. *Tetrahedron Lett.* **1977**, 3325.

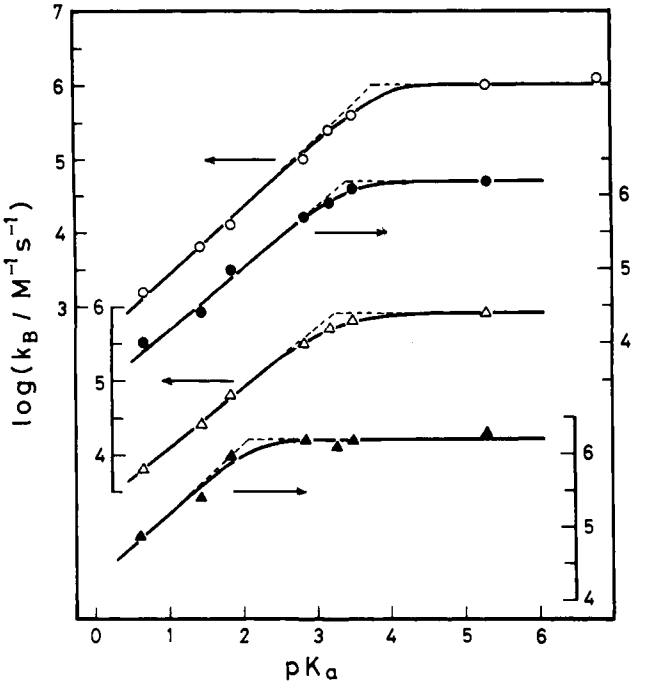


Figure 4. Brønsted plots of log k_B for the proton-transfer reactions from the radical cations of PyH_2 [BNAH (○), 4-ClBNAH (●), 2,4-Cl_2BNAH (△), AcH_2 (▲)] to pyridine derivatives in MeCN at 298 K vs. pK_a of pyridine derivatives.

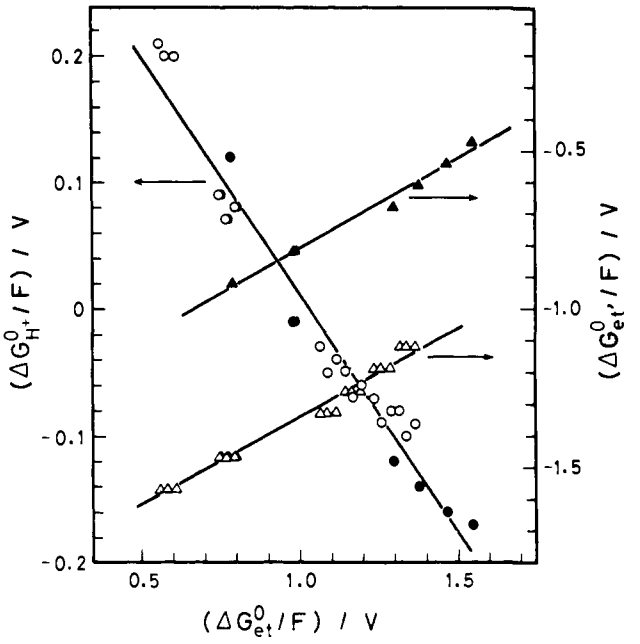


Figure 5. Correlations of the Gibbs energy changes of the proton-transfer reactions from X-BNAH$^{\bullet+}$ (○) and $AcH_2^{\bullet+}$ (●) to $Q^{\bullet-}$ ($\Delta G°_{H^+}$) and the electron-transfer reactions from X-BNA$^{\bullet}$ (△) and AcH$^{\bullet}$ (▲) to QH$^{\bullet}$ ($\Delta G°_{et'}$) with the Gibbs energy change of the electron-transfer reactions from X-BNAH and AcH_2 to Q ($\Delta G°_{et}$).

equal to the initial concentration of PyH_2. When $Mg(ClO_4)_2$ was added to the PyH_2–Q system, only the hydride-transfer reaction (eq 15) occurred, and the radical anion $Q^{\bullet-}$ has not been formed as confirmed by the electronic spectrum.[21] Rates of the hydride-transfer reactions from PyH_2 to Q in the presence of Mg^{2+} ion in MeCN were followed by the decay of the PyH_2 concentration, which is expressed by eq 18. In the case of hydride-

$$-d[PyH_2]/dt = k_{obsd}[PyH_2][Q] \quad (18)$$

transfer reactions from AcH_2 to Q, the rates were followed by the formation of AcH^+. The observed rate constants k_{obsd} were determined under pseudo-first-order conditions in the presence of a large excess Q. The k_{obsd} values in the absence and presence of 0.10 M Mg^{2+} ion are summarized in Table VI, together with the one-electron reduction potentials of Q in the absence[20] and presence of 0.10 M Mg^{2+} ion in MeCN.[21] It is shown in Table VI that the rate constant in the absence of Mg^{2+} ion increases with a positive shift of the reduction potential of Q $E°(Q/Q^{\bullet-})$, spanning a range of more than 10^{11}, and the rate constant with the same quinone decreases in the order 4-MeOBNAH > 4-MeBNAH > BNAH > 4-ClBNAH > 2,4-Cl_2BNAH > AcH_2 in accordance with an increase of the $E°(PyH_2^{\bullet+}/PyH_2)$ value (Table VI).

Discussion

Gibbs Energy Changes of Electron, Proton, Hydrogen, and Hydride Transfers. On the basis of the results in the present study, we can now evaluate the Gibbs energy change of each step of the hydride-transfer reaction from PyH_2 to Q, i.e., that of the initial electron transfer from PyH_2 to Q ($\Delta G°_{et}$), the subsequent proton transfer from $PyH_2^{\bullet+}$ to $Q^{\bullet-}$ ($\Delta G°_{H^+}$), and the second electron transfer from PyH$^{\bullet}$ to QH$^{\bullet}$ ($\Delta G°_{et'}$) as well as the Gibbs energy change of the hydrogen transfer from PyH_2 to Q ($\Delta G°_{H}$) and that of the overall transfer of a hydride ion from PyH_2 to Q ($\Delta G°_{H^-}$) as follows. The $\Delta G°_{et}$ values are obtained from the one-electron redox potentials of PyH_2 $E°(PyH_2^{\bullet+}/PyH_2)$ and Q $E°(Q/Q^{\bullet-})$ by using eq 19

$$\Delta G°_{et}/F = E°(PyH_2^{\bullet+}/PyH_2) - E°(Q/Q^{\bullet-}) \quad (19)$$

The $E°(PyH_2^{\bullet+}/PyH_2)$ and $E°(Q/Q^{\bullet-})$ values are given in Table III and VI, respectively. The $\Delta G°_{et'}$ values are also obtained from the one-electron redox potentials of PyH^+ $E°(PyH^+/PyH^{\bullet})$ and $QH^{\bullet}$ $E°(QH^{\bullet}/QH^-)$ by using eq 20 where the $E°(PyH^+/PyH^{\bullet})$

$$\Delta G°_{et'}/F = E°(PyH^+/PyH^{\bullet}) - E°(QH^{\bullet}/QH^-) \quad (20)$$

values are also given in Table III, and the $E°(QH^{\bullet}/QH^-)$ values of various *p*-benzoquinone derivatives have been reported by Rich et al.[25a] On the other hand, the $\Delta G°_{H^+}$ values can be determined from the pK_a values of $PyH_2^{\bullet+}$ and QH$^{\bullet}$ by using eq 21 where

$$\Delta G°_{H^+} = 2.3RT[pK_a(PyH_2^{\bullet+}/PyH^{\bullet}) - pK_a(QH^{\bullet}/Q^{\bullet-})] \quad (21)$$

the pK_a($PyH_2^{\bullet+}$/PyH$^{\bullet}$) values are listed in Table V, and the pK_a($QH^{\bullet}/Q^{\bullet-}$) values are given in the literatures.[25a] It should be noted that the $\Delta G°_{et'}$ and $\Delta G°_{H^+}$ values must be accepted with reservation since the reported values of $E°(QH^{\bullet}/QH^-$, and pK_a($QH^{\bullet}/Q^{\bullet-}$)[25a] in eq 10 and 21, respectively, are those in H_2O, and the use as the values in MeCN may not be correct in an absolute sense becuase of the different solvation between H_2O and MeCN.[54] However, these values can be safely used for comparison between different NADH model compounds (X-BNAH and AcH_2), which is the main theme of this study.[55]

By using these values obtained by eq 19–21, the correlations of $\Delta G°_{et}$ with $\Delta G°_{H^+}$ and $\Delta G°_{et'}$ are obtained as shown in Figure 5.[56] It can be seen that there is a single linear correlation between $\Delta G°_{H^+}$ and $\Delta G°_{et}$ for both X-BNAH (open circles) and AcH_2 (closed circles) as expressed by eq 22 but that there are two

$$\Delta G°_{H^+} = 0.39F - 0.37\Delta G°_{et} \text{ (for both X-BNAH and } AcH_2) \quad (22)$$

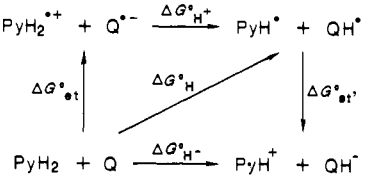
separate linear correlations between $\Delta G°_{et'}$ and $\Delta G°_{et}$ for X-

(54) However, the sensitivity of $E°(Q/Q^{\bullet-})$ to substituent effects is approximately independent of solvents (H_2O and MeCN); for methyl-substituted *p*-benzoquinone derivatives, $E°(Q/Q^{\bullet-}$ in $H_2O)^{25a} = 0.36 + 1.0E°(Q/Q^{\bullet-}$ in MeCN),[20a] where the correlation coefficient is 0.999.

(55) Both the one-electron oxidation potentials of X-BNAH and AcH_2 as well as the one-electron reduction potentials of X-BNA$^+$ and AcH$^+$ have been determined in the same solvent (i.e., MeCN, see Table III).

(56) The plots in Figure 5 include the data of PyH_2 (BNAH, 4-ClBNAH, 2,4-Cl_2BNAH, AcH_2) for which all the values of $E°(PyH_2^{\bullet+}/PyH_2)$, $E°(PyH^+/PyH^{\bullet})$, and pK_a($PyH_2^{\bullet+}$/PyH$^{\bullet}$) have been determined in this study and those Q (*p*-chloranil, 2,6-dichloro-*p*-benzoquinone, 2,5-dichloro-*p*-benzoquinone, *p*-benzoquinone, methyl-*p*-benzoquinone, 2,6-dimethyl-*p*-benzoquinone, trimethyl-*p*-benzoquinone) for which all the values of $E°(Q/Q^{\bullet-})$, $E°(QH^{\bullet}/QH^-)$, and pK_a($QH^{\bullet}/Q^{\bullet-}$) are known.

Scheme II



BNAH (open triangles) and AcH_2 (closed triangles) as expressed by eq 23 and 24, respectively. Such different intercepts between

$$\Delta G°_{et'} = -1.91F + 0.56\Delta G°_{et} \quad \text{(for X-BNAH)} \quad (23)$$

$$\Delta G°_{et'} = -1.38F + 0.57\Delta G°_{et} \quad \text{(for } AcH_2\text{)} \quad (24)$$

X-BNAH and AcH_2 with the same slope (eq 23 and 24) are ascribed to the large difference in the one-electron reduction potentials between X-BNA^+ (−1.08 V) and AcH^+ (−0.43 V).

The relationship of $\Delta G°_H$ and $\Delta G°_{H^-}$ with $\Delta G°_{et}$, $\Delta G°_H$, and $\Delta G°_{et'}$ is shown in Scheme II. It is readily recognized that $\Delta G°_H$ and $\Delta G°_{H^-}$ are given by eq 25 and 26, respectively. Thus, by

$$\Delta G°_H = \Delta G°_{et} + \Delta G°_{H^+} \quad (25)$$

$$\Delta G°_{H^-} = \Delta G°_{et} + \Delta G°_{H^+} + \Delta G°_{et'} \quad (26)$$

combining eq 22–24 with eq 25 and 26, $\Delta G°_H$ is expressed by the same linear function of $\Delta G°_{et}$ for both X-BNAH and AcH_2 while $\Delta G°_{H^-}$ gives two linear correlations with $\Delta G°_{et}$ between X-BNAH andAcH_2 with different intercepts albeit the slope is the same, since $\Delta G°_{et'}$ is different between X-BNAH and AcH_2 owing to the large difference in the one-electron reduction potentials between X-BNA^+ and AcH^+ (Table III).

Gibbs Energy Relationships of Hydride-Transfer Reactions. The logarithms of rate constants for the hydride-transfer reactions from both X-BNAH and AcH_2 to Q in the absence of Mg^{2+} ion in MeCN are plotted against $\Delta G°_{H^-}/F$ and $\Delta G°_H/F$ as shown in Figure 6 (parts a and b, respectively,[57] where log k_{obsd} is expressed as two different linear functions of $\Delta G°_{H^-}$ between X-BNAH and AcH_2 (eq 27 and 28 respectively),[58] while log k_{obsd} is given by the same linear function of $\Delta G°_H$ for both X-BNAH and AcH_2 (eq 29).[58] The slopes of the plots of log k_{obsd} vs. $\Delta G°_{H^-}/F$ (eq 27

$$\log k_{obsd} = -3.5 - 7.9(\Delta G°_{H^-}/F) \quad \text{(for X-BNAH)} \quad (27)$$

$$\log k_{obsd} = 0.82 - 8.5(\Delta G°_{H^-}/F) \quad \text{(for } AcH_2\text{)} \quad (28)$$

$$\log k_{obsd} = 14.6 - 15.3(\Delta G°_H/F) \quad \text{(for X-BNAH and } AcH_2\text{)} \quad (29)$$

and 28) show that as the reaction becomes more exothermic, about one-half of the change in $\Delta G°_{H^-}$ is reflected in the activation barrier of the hydride-transfer reactions, i.e., the Brønsted slope $\alpha = 0.5$.[59] A similar Brønsted value has been reported for hydride-transfer reactions from NADH to *p*-benzoquinone and *o*-benzoquinone derivatives, being considered to support a direct hydride-transfer mechanism although it has been pointed out that a hydrogen-transfer mechanism cannot be ruled out.[22] However, if the intrinsic barrier of the reaction (i.e., the activation Gibbs energy when the Gibbs energy change is zero) is sufficiently large, any homologous series of reactions gives a linear Gibbs energy relationship between log k_{obsd} and the Gibbs energy change of the reaction with $\alpha = 0.5$,[60] providing no definitive information on the nature of the transition states. Thus, a direct hydride-transfer mechanism can provide no quantitative interpretation for the different correlation of log k_{obsd} with $\Delta G°_{H^-}$ between X-BNAH and AcH_2.

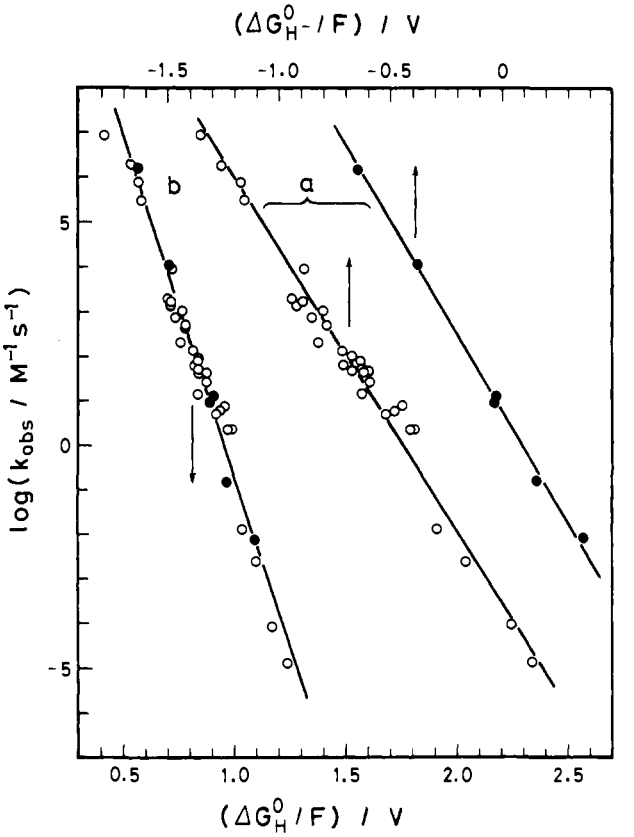


Figure 6. Plots of log k_{obsd} for the hydride-transfer reactions from PyH_2 [X-BNAH (○) and AcH_2 (●)] to Q in MeCN at 298 K vs. (a) the Gibbs energy change of the hydride-transfer reactions from PyH_2 to Q ($\Delta G°_{H^-}/F$) and (b) the Gibbs energy change of the hydrogen-transfer reactions from PyH_2 to Q ($\Delta G°_H/F$). The linear solid lines are drawn by the least-squares curve fits.[59]

On the other hand, the slope of *single* correlation of log k_{obsd} with $\Delta G°_H$ (eq 29), which is about twice larger than the slopes in the plots of log k_{obsd} vs. $\Delta G°_{H^-}/F$ (eq 22 and 28) indicates that the change in $\Delta G°_H$ is directly reflected in the activation barrier of the hydride-transfer reactions, i.e., α is close to unity (0.91). Thus, the *single* correlation of log k_{obsd} with $\Delta G°_H$ for both X-BNAH and AcH_2, combined with the different intercepts between X-BNAH and AcH_2 in the plots of log k_{obsd} vs. $\Delta G°_{H^-}$ (= $\Delta G°_H + \Delta G°_{et'}$), reveals that the activation barrier of the hydride-transfer reactions from PyH_2 to Q is dependent on only $\Delta G°_H$ and thereby independent of $\Delta G°_{et'}$; the different correlation of log k_{obsd} with $\Delta G°_{H^-}$ between X-BNAH and AcH_2 is caused mainly by the difference of the reduction potentials between X-$BNA^•$ (−1.08 V) and $AcH^•$ (−0.43 V), which has no effect on the activation barrier of the overall hydride transfer.[61] The apparent correlation of log k_{obsd} with $\Delta G°_{H^-}$ may hold only provided that $\Delta G°_{et'}$ is correlated with $\Delta G°_H$ in a homologous series of reactions. Since $\Delta G°_H$ consists of $\Delta G°_{et}$ and $\Delta G°_{H^+}$ (eq 25), the mechanistic question is reduced to (a) a direct transfer of a hydrogen atom from PyH_2 to Q, followed by an exothermic electron transfer from $PyH^•$ to $QH^•$ or (b) an electron transfer from PyH_2 to Q in the activation process and the subsequent proton transfer from $PyH_2^{•+}$ to $Q^{•-}$, followed by an exothermic electron transfer form $PyH^•$ to $QH^•$.

(57) In the plots in Figure 6, the $\Delta G°_H$ and $\Delta G°_{H^-}$ values of the PyH_2–Q systems, which are not included in Figure 5, are obtained by using eq 25 and 26, respectively, where the $\Delta G°_{H^+}$ and $\Delta G°_{et'}$ values are estimated by using eq 22 and 23 (or 24), respectively.

(58) The linear solid lines in Figure 6 are drawn by the least-squares curve fits by using the data in the region where the linear relations of $\Delta G°_{H^+}$ and $\Delta G°_{et'}$ with $\Delta G°_{et}$ are shown to hold (Figure 5).

(59) The slope of the plots in Figure 6a corresponds to $-\alpha F/2.3RT$ and, therefore, $\alpha = 8.4 \times 2.3RT/F = 0.50$.

(60) (a) Glasstone, S.; Laidler, K. J.; Eyring, H. *The Theory of Rate Processes*; McGraw Hill: New York, 1941; p 145. (b) Levine, R. D. *J. Phys. Chem.* **1979**, *83*, 159.

(61) The differences in the $\Delta G°_{H^-}/F$ value between BNAH and AcH_2 in H_2O ($\Delta\Delta G°_{H^-}/F$) has been reported to be −0.57 V (ref 11a, see, also: Ostovic, D.; Lee, I.-S. H.; Roberts, R. M. G.; Kreevoy, M. M. *J. Org. Chem.* **1985**, *50*, 4206). The $\Delta\Delta G°_{H^2}/F$ value in MeCN determined in this study is −0.84 V, which is more negative than the value in H_2O. This difference between MeCN and H_2O may be ascribed to the more positive one-electron oxidation potential of the dihydronicotinamide in H_2O than that in MeCN (see ref 44).

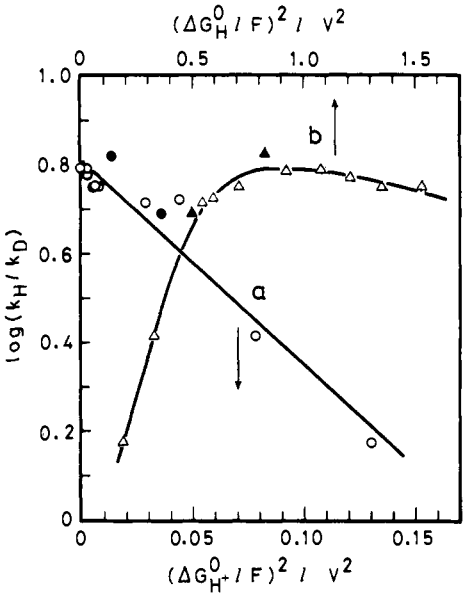


Figure 7. Plots of log (k_H/k_D) for the hydride-transfer reactions from BNAH[20] and AcH_2[53a] to Q in MeCN vs. (a) $(\Delta G°_{H^+}/F)^2$ [BNAH (○), AcH_2 (●)] and (b) $(\Delta G°_H/F)^2$ [BNAH (△), AcH_2 (▲)], see eq 31 and 32, respectively.

Kinetic Isotope Effects. The kinetic isotope effects of hydride-transfer reactions from PyH_2 to Q may provide useful information to determine whether an actual transfer of hydrogen nucleus occurs as a form of proton or hydrogen atom. In the case of the electron-transfer activation mechanism, the observed primary kinetic isotope effects k_H/k_D may be attributed to those for the proton transfer from $PyH_2^{\bullet+}$ to $Q^{\bullet-}$ since no kinetic isotope effect is expected for the electron-transfer process from PyH_2 to Q as well as $PyH^\bullet$ to $QH^\bullet$. In such a case, k_H/k_D may be expressed as a function of the Gibbs energy change of the proton-transfer $\Delta G°_{H^+}$, by using the Marcus equation (eq 30),[50d,62] where

$$\Delta G^{\ddagger}_{H^+} = \frac{\lambda_{H^+}}{4}\left(1 + \frac{\Delta G°_{H^+}}{\lambda_{H^+}}\right)^2 \qquad (30)$$

λ_{H^+} is the reorganization energy for the proton transfer and $\lambda_{H^+}/4$ corresponds to the activation Gibbs energy when $\Delta G°_{H^+} = 0$. From eq 30 is derived eq 31, where $\log(k_H/k_D)_0$ is the maximum

$$\log\left(\frac{k_H}{k_D}\right) = \log\left(\frac{k_H}{k_D}\right)_0\left(1 - \frac{(\Delta G°_{H^+})^2}{\lambda_{H^+}\lambda_{D^+}}\right) \qquad (31)$$

kinetic isotope effect which is equal to $(\lambda_{D^+} - \lambda_{H^+})/(4 \times 2.3RT)$. If an actual transfer of hydrogen nucleus occurs as a form of hydrogen atom, eq 32 would be applied instead of eq 31, where $\Delta G°_{H^+}$, λ_{H^+}, and λ_{D^+} are substituted by $\Delta G°_H$, λ_H, and λ_D, respectively.[63]

$$\log\left(\frac{k_H}{k_D}\right) = \log\left(\frac{k_H}{k_D}\right)_0\left(1 - \frac{(\Delta G°_H)^2}{\lambda_H\lambda_D}\right) \qquad (32)$$

The reliable k_H/k_D values have been reported for the hydride-transfer reactions from BNAH[20] and AcH_2[53] to Q by using [4,4-D_2]BNAH and [9,9-D_2]AcH_2, respectively, in MeCN.[64] The $\log(k_H/k_D)$ values are plotted against the $(\Delta G°_{H^+})^2$ values as shown in Figure 7a, where $\log(k_H/k_D)$ decreases linearly with increasing the $(\Delta G°_{H^+}/F)^2$ value in accordance with eq 31. From the slope and the intercept of the plot in Figure 7a, the λ_{H^+} and λ_{D^+} values are determined as 7.8 and 12.2 kcal mol^{-1}, respectively. On the other hand, no linear correlation is observed between $\log(k_H/k_D)$ and $(\Delta G°_H/F)^2$ as shown in Figure 7b, indicating that the primary kinetic isotope effect is not correlated with the Gibbs energy change of the hydrogen transfer $\Delta G°_H$. The variation of k_H/k_D in Figure 7, which shows a sharp contrast to approximately constant k_H/k_D values (4.42–5.95) in the hydride-transfer reactions between various NADH and NAD^+ analogues,[11b,65] cannot be related to the Gibbs energy change of the overall hydride transfer $\Delta G°_{H^-}$, either. Such a variation of the primary kinetic isotope effect is inconsistent with the mechanism of one-step hydride transfer since if an actual transfer of hydrogen nucleus occurs as a form of hydride ion, the constant Brønsted slope ($\alpha = 0.5$) in Figure 5a (eq 27 and 28) would require the constant k_H/k_D value with the change in $\Delta G°_{H^-}$. Thus, it is concluded that an actual transfer of hydrogen nucleus occurs as a form of proton and the electron-transfer activation mechanism seems to be the most plausible one to explain both the variation of the rate constants and the primary kinetic isotope effects in the hydride-transfer reactions from PyH_2 to Q.

In the next section, we wish to compare between the hydride-transfer and electron-transfer reactions of NADH model compounds in order to clarify the relation between them.

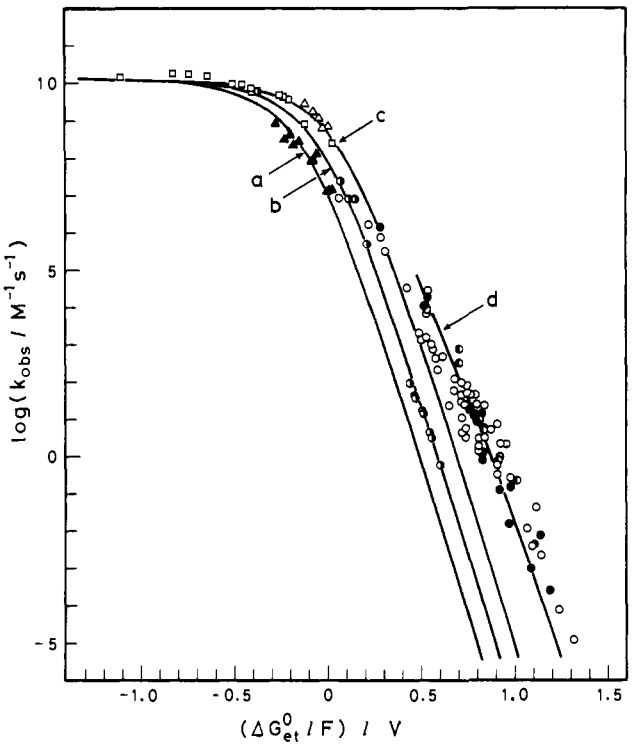


Figure 8. Plots of log k_{obsd} vs. the Gibbs energy change of the electron-transfer reactions of NADH model compounds for (a) the electron-transfer reactions from X-BNAH to $[RuL_3]^{2+*}$ in the absence and presence of Mg^{2+} ion (1.0×10^{-2} M) in MeCN (▲), (b) those from NADH to ferricenium ions[44] and organic monocations[66] in H_2O (◑), (c) those from X-BNAH to $[RuL_2L']^*$ in MeCN (△) and those from BNAH* to organic oxidants in the absence and presence of Mg^{2+} ion (5.0×10^{-2} M) in MeCN (□), and (d) the hydride-transfer reactions from X-BNAH (○) and AcH_2 (●) to Q in the absence and presence of Mg^{2+} ion (0.10 M) in MeCN at 298 K and those from NADH to Q in H_2O at 303 K[22] (◐). The solid lines in Figure 8a–c are drawn based on eq 35 by using the work terms w_p −0.10, 0, +0.10 eV, respectively, and the simulated solid line in Figure 8d is drawn based on eq 39 by using the work term −0.31 eV, see text.

(62) Marcus, R. A. *J. Phys. Chem.* **1968**, *72*, 891.

(63) The kinetic isotope effects on hydrogen-transfer reactions have been reported to show a maximum when the heat of reaction ΔH is approximately zero and decreases with an increase in the $|\Delta H|$ value, see: Pryor, W. A.; Kneipp, K. G. *J. Am. Chem. Soc.* **1971**, *93*, 5584.

(64) It has been pointed out that the k_H/k_D values determined by using [4-D_1]BNAH instead of [4,4-D_2]BNAH may include large experimental errors (see ref 9a).

(65) Although the k_H/k_D values have been reported to be constant with the change in the equilibrium constant for the hydride-transfer reactions between NADH and NAD^+ analogues (K_{H^-}) in the region $0.70 < K_{H^-} < 2 \times 10^6$ M^{-1} (ref 11b), the corresponding variation of $\Delta G°_{H^-}/F$ (−0.01 ~ 0.37 V), which is much smaller than that in this study (Figure 7), may not be wide enough to provide a definitive proof that an actual transfer of hydrogen nucleus occurs as a form of hydride ion.

Relation between Electron-Transfer and Hydride-Transfer Reactions of NADH and Its Model Compounds. Extensive comparison between the hydride-transfer and electron-transfer reactions of NADH and its model compounds in the absence and presence of Mg^{2+} ion is shown in Figure 8, where the logarithms of rate constants are plotted against the standard Gibbs energy change of the electron transfer $\Delta G°_{et}/F$. Electron transfer reactions of NADH and its model compounds are classified into three groups by using differently charged oxidants (neutral, monocation, and dication); photoinduced electron-transfer reactions of X-BNAH with various neutral organic oxidants (Table II) and $[RuL_2L']$ (Table I) in MeCN, electron-transfer reactions of NADH with organic monocations[66] and ferricenium cations (Fc^+)[44] in H_2O, and photoinduced electron-transfer reactions of X-BNAH with a dication complex $[RuL_3]^{2+}$ in MeCN (Table I) are shown in Figure 8 (parts a, b, and c, respectively). The $\Delta G°_{et}$ values are obtained from the one-electron oxidation potentials of reductants $E°(red^{\bullet+}/red)$ (X-BNAH in Table III, NADH,[43] and BNAH*[67]) and the one-electron reduction potentials of oxidants $E°(ox/ox^{\bullet-})$ ($[RuL_3]^{2+*}$,[68] $[RuL_2L']^*$,[69] and organic oxidants in Table II) by using eq 33.

$$\Delta G°_{et}/F = E°(red^{\bullet+}/red) - E°(ox/ox^{\bullet-}) \quad (33)$$

By applying the general scheme for the fluorescence quenching by an outersphere electron transfer in MeCN to the present system (Scheme III),[70,71] the observed rate constant k_{obsd} may be expressed

Scheme III

$$red^* + ox \underset{k_{21}}{\overset{k_{12}}{\rightleftharpoons}} [red^*\ ox] \underset{k_{32}}{\overset{k_{23}}{\rightleftharpoons}} [red^+\ ox^-] \xrightarrow{k_{30}} \text{decay}$$

by eq 34, where k_{30} comprises all possible modes by which the

$$k_{obsd} = \frac{k_{12}}{1 + \frac{k_{21}}{k_{30}}\left(\frac{k_{30}}{k_{23}} + \frac{k_{32}}{k_{23}}\right)} \quad (34)$$

radical ion pair disappears, in particular via the back electron transfer to the ground state, being approximately equal to the frequency factor. Under such a condition, eq 34 is rewritten by eq 35,[70] where $\Delta G_{23}^{\ddagger}$ and ΔG_{23} are the activation Gibbs energy

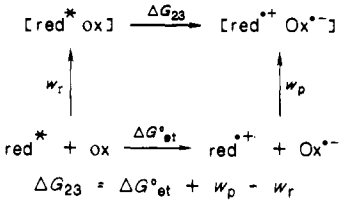
$$k_{obsd} = \frac{2.0 \times 10^{10}}{1 + 0.25[\exp(\Delta G_{23}^{\ddagger}/RT) + \exp(\Delta G_{23}/RT)]} \quad (35)$$

and the Gibbs energy change of the actual electron-transfer process, $[red^*\ ox] \rightarrow [red^{\bullet+}\ ox^{\bullet-}]$, respectively. The relation between $\Delta G_{23}^{\ddagger}$ and ΔG_{23} is well-expressed by the Rehm–Weller equation (eq 36),[70] where $\Delta G_0^{\ddagger}$ is the activation Gibbs energy when

$$\Delta G_{23}^{\ddagger} = (\Delta G_{23}/2) + [(\Delta G_{23}/2)^2 + (\Delta G_0^{\ddagger})^2]^{1/2} \quad (36)$$

$\Delta G_{23} = 0$. On the other hand, the relation between ΔG_{23} and $\Delta G°_{et}$ is shown in Scheme IV and given by eq 37 where w_p and w_r are the work terms required to bring the products ($red^{\bullet\ddagger}$ and $ox^{\bullet-}$) and the reactants (red* and ox) together to the mean separation in the activated complex, which are largely coulombic, and w_r may be neglected since the reactants in the present case include neutral species. Thus, the k_{obsd} value can be calculated as a function of $\Delta G°_{et}$ by using eq 35–37.

Scheme IV



$$\Delta G_{23} = \Delta G°_{et} + w_p - w_r \quad (37)$$

The calculated dependence of log k_{obsd} on $\Delta G°_{et}$ by assuming that $w_p = 0$ and $\Delta G_0^{\ddagger} = 4.0$ kcal mol^{-1} agrees well with the experimental results for the electron-transfer reactions of NADH with organic monocations[66] and ferricenium ions (Fc^+)[44] as shown by the solid line in Figure 8b. Such an agreement indicates that the work term w_p is in fact negligible when the products upon the electron transfer include a neutral species such as Fc. On the other hand, when the calculated curve of log k_{obsd} vs. $\Delta G°_{et}/F$ is shifted to −0.1 and +0.1 V in the abscissa, the shifted curves agree well with the experimental results for the electron-transfer reactions between X-BNAH and a dication complex $[RuL_3]^{2+*}$ and those for the reactions between neutral reactants (the BNAH*-organic oxidant and X-BNAH-$[RuL_2L']^*$ systems) as shown by the solid lines in Figure 8 (parts a and 8c, respectively). Such agreements indicate that the w_p/F values of the radical ion pairs of the like charges (X-$BNAH^{\bullet+}$) $[RuL_3]^+$) and the opposite charges [($BNAH^{\bullet+}$ $ox^{\bullet-}$) and (X-$BNAH^{\bullet+}$ $[RuL_2L']^-$)] are +0.1 and −0.1 eV, respectively.[72]

The plots in Figure 8a and c include the data in the presence of Mg^{2+} ion, which agree with the data in the absence of Mg^{2+} ion in the calculated dependence of log k_{obsd} on $\Delta G°_{et}/F$. Thus, the retarding effect of Mg^{2+} ion on the photoinduced electron-transfer reactions of X-BNAH (Table I and II) corresponds to the positive shift of $\Delta G°_{et}$ in the presence of Mg^{2+} ion, which is attributed to the positive shift of the one-electron oxidation potentials of the ground and excited states of X-BNAH (Table III) due to the complex formation with Mg^{2+} ion (eq 4).

In Figure 8d, the log k_{obsd} values for the hydride transfer reactions from PyH_2 (X-BNAH, AcH_2, NADH[22]) to a series of *p*-benzoquinone derivatives Q in the absence and presence of Mg^{2+} ion (0.10 M) are plotted against the Gibbs energy change of the electron transfer from PyH_2 to Q, $\Delta G°_{et}/F$, which is obtained by eq 19. An important point to note in Figure 8d is that there is an approximately *single* correlation between log k_{obsd} and $\Delta G°_{et}/F$ for different kinds of dihydropyridine compounds (X-BNAH, AcH_2, and NADH shown by the open, closed, and half-closed circles, respectively) in the absence and presence of Mg^{2+} ion, indicating that the activation barrier of the hydride-transfer reactions is well-correlated with the energetics of the electron transfer $\Delta G°_{et}$. However, when the plot in Figure 8d is compared with the calculated dependence of log k_{obsd} on $\Delta G°_{et}/F$ for the electron-transfer reactions of X-BNAH (the solid line in Figure 8c), the rate constants for the hydride-transfer reactions are greater than those of the electron-transfer reactions, although they become closer to each other with a decrease in $\Delta G°_{et}$. Such a difference in the reactivity has been taken to argue against the involvement of an electron-transfer process in the hydride-transfer reactions by assuming that the work term w_p of the radical ion pair produced upon electron transfer can be neglected.[22,73] However, the work

(66) Grodkowski, J.; Neta, P.; Carlson, B. W.; Miller, L. L. *J. Phys. Chem.* **1983**, *87*, 3135.

(67) The one-electron oxidation potential of BNAH* is obtained at −2.60 V by subtracting the zero–zero transition energy $\Delta E_{0,0}$ (= 3.17 eV) from $E°(BNAH^{\bullet+}/BNAH)$ (= 0.57 V).[38]

(68) The one-electron reduction potential of $[RuL_3]^{2+*}$ has been reported to be 0.77 V (ref 36a).

(69) The one-electron reduction potential of $[RuL_2L']^*$ is obtained as 0.61 V by adding the zero–zero transition energy $\Delta E_{0,0}$ (= 2.00 eV) to $E°([RuL_2L']/[RuL_2L']^-)$ (= −1.39 V), which was determined by the cyclic voltammetry (see Experimental Section).

(70) (a) Rehm, D.; Weller, A. *Isr. J. Chem.* **1970**, *8*, 259. (b) Rehm, D.; Weller, A. *Bunsenges. Phys. Chem.* **1969**, *73*, 834.

(71) The asterisk identifies the excited state, and Scheme IV is applied to the electron-transfer reactions between BNAH* and organic oxidants. In the case of the electron-transfer reactions between PyH_2 and $[RuL_3]^{2+*}$ or $[RuL_2L']^*$, the asterisk should be put on ox instead of red.

(72) Similar work terms w_p for the fluorescence quenching of some aromatic acceptors by neutral donors and the luminescence quenching of $[RuL_3]^{2+*}$ by neutral donors have been reported to be −0.12 and +0.13 eV, respectively.[38]

(73) Similar discussion has recently appeared in ref 23, where the redox potentials in MeCN are used to discuss the rates in H_2O. It should be pointed out, however, that the difference in $E°(Q/Q^{\bullet-})$ between H_2O and MeCN (ref 54) may correspond to the difference of 10^6 in the rate. Thus, the discussion on the absolute values of rate constants in relation with the redox potentials must be made by using the same solvent as shown in Figure 8.

term w_p of the radical ion pair of the opposite charges produced by innersphere electron-transfer reactions via the charge-transfer (CT) complexes is expected to be much more negative than w_p in the case of outersphere electron-transfer reactions,[2] such as photoinduced electron-transfer reactions of X-BNAH with organic oxidants and $[RuL_2L']$ (Figure 8c) in which the w_p value is obtained as −2.3 kcal mol^{-1} (w_p/F = −0.1 eV) as described above. In fact, the w_p value of the radical ion pair for the reaction of BNAH with 2,3-dichloro-5,6-dicyano-*p*-benzoquinone has previously been evaluated as −7.1 kcal mol^{-1} (w_p/F = −0.31 eV), which is much more negative than the w_p value of the outersphere electron-transfer reactions, from the formation constant of the radical ion pair (log K_{et} = 4.3),[20] by using eq 38,[74] where the $E°$(BNAH$^{•+}$/BNAH) and $E°$(Q/Q$^{•-}$) values are given in Table III and VI, respectively.

$$2.3RT \log K_{et} = F[E°(\mathrm{BNAH}^{•+}/\mathrm{BNAH}) - E°(\mathrm{Q}/\mathrm{Q}^{•-})] - w_p \quad (38)$$

In the following discussion, we wish to show that the difference in the reactivity between the electron-transfer and hydride-transfer reactions (Figure 8c and d, respectively) can be analyzed quantitatively by using the w_p value (−7.1 kcal mol^{-1}), based on the electron-transfer activation mechanism for the hydride-transfer reactions from PyH_2 to Q (Scheme V). Scheme V is essentially

Scheme V

$$\mathrm{PyH_2} + \mathrm{Q} \underset{k_{21}}{\overset{k_{12}}{\rightleftarrows}} [\mathrm{PyH_2\ Q}] \underset{k_{32}}{\overset{k_{23}}{\rightleftarrows}} [\mathrm{PyH_2}^{•+}\ \mathrm{Q}^{•-}] \xrightarrow{k_{H^+}} [\mathrm{PyH}^{•}\ \mathrm{QH}^{•}] \xrightarrow{\text{fast}} \mathrm{PyH}^{+} + \mathrm{QH}^{-}$$

the same as Scheme III, except for the proton-transfer process (k_{H^+}) being the route of the disappearance of the radical ion pair instead of the back electron-transfer process to the ground state (k_{30}) in Scheme III, when the asterisk which denotes the excited state is removed from the reductant in Scheme III. Then, the observed rate constant can be expressed by eq 39 by a similar

$$k_{obsd} = \frac{2.0 \times 10^{10}}{1 + 0.25\left[\exp\left(\frac{\Delta G_{23}^{\ddagger}}{RT}\right) + \exp\left(\frac{\Delta G^{\ddagger}_{H^+} + \Delta G_{23}}{RT}\right)\right]} \quad (39)$$

manner to derive eq 35. The relation between $\Delta G_{23}^{\ddagger}$ and ΔG_{23} is given by eq 36, and ΔG_{23} is given by eq 37, where the w_p and w_r values are taken as −7.1 kcal mol^{-1} and 0, respectively. On the other hand, $\Delta G^{\ddagger}_{H^+}$ is given as a function of $\Delta G°_{H^+}$ (eq 30) which can be obtained from $\Delta G°_{et}$ by eq 22. Thus, by combining together, the k_{obsd} value is calculated as a function of $\Delta G°_{et}$ by using eq 39. The calculated results are illustrated by the solid line in Figure 8d, which shows reasonable agreements with the experimental results for the hydride-transfer reactions from different kinds of dihydropyridine compounds (X-BNAH, AcH_2, NADH[22]) to a series of *p*-benzoquinone derivatives Q in both the absence and presence of Mg^{2+} ion.[75] Such agreements strongly support the electron-transfer activation mechanism of the hydride-transfer reactions from PyH_2 to Q (Scheme V).[76] Since the data in the absence of Mg^{2+} ion agree well with those in the presence of Mg^{2+} ion in the calculated dependence of log k_{obsd} on $\Delta G°_{et}$, the retarding and accelerating effects of Mg^{2+} ion on the hydride-transfer reactions depending on Q (Table VI) correspond to the shifts of $\Delta G°_{et}$ in the presence of Mg^{2+} ion, which are attributed to the positive shifts of the one-electron oxidation potentials of X-BNAH (Table III) as well as the one-electron reduction potentials of Q due to the complex formation with Mg^{2+} ion,[21] according to eq 33.

Summary and Conclusion

The electron-transfer activation mechanism of the hydride-transfer reactions from PyH_2 to Q (Scheme V) has been shown to provide a much more quantitative description of the energetic profiles of the reactions than the mechanism of a direct transfer of a hydride ion or hydrogen atom by the following reasons. (1) The log k_{obsd} values of the hydride-transfer reactions can be expressed as the *same* function of the Gibbs energy change of the electron transfer from PyH_2 to Q ($\Delta G°_{et}$) for both X-BNAH and AcH_2 (Figure 8d), which is successfully simulated based on Scheme V, but the log k_{obsd} values are expressed as two *different* correlations between X-BNAH and AcH_2 with the Gibbs energy change of the overall hydride transfer ($\Delta G°_{H^-}$) (Figure 6a). (2) The variation of the primary kinetic isotope effect (Figure 7a), which can be predicted quantitatively based on Scheme V, indicated that an actual transfer of hydrogen nucleus occurs as a form of proton. Neither a direct transfer of a hydride ion nor hydrogen atom can provide a reasonable explanation on the variation of the kinetic isotope effect, which must be consistent with the variation of the rate constant. (3) The retarding and accelerating effects of Mg^{2+} ion on the hydride-transfer reactions depending on the substrate Q can be simulated quantitatively based on Scheme V (Figure 8d) by the positive shifts of the one-electron redox potentials of X-BNAH and Q, while neither mechanism of a direct transfer of a hydride ion nor hydrogen atom provides any quantitative explanation on the dual effects of Mg^{2+} ion. We have recently shown that the electron-transfer activation mechanism (Scheme V) can provide the quantitative basis also on the retarding and accelerating effects of protons on the hydride-transfer reactions from AcH_2 to Q depending on Q.[77] (4) The difference in the reactivity between the photoinduced electron-transfer reactions and the hydride-transfer reactions via the electron-transfer activation (Scheme V) is clarified as being caused by the difference in the work term w_p required to bring the product ions to the mean separation of the activated complex between the outersphere electron-transfer and the innersphere electron-transfer reactions, respectively.[76]

Notwithstanding the similarities between the electron-transfer reactions (Scheme III) and the electron-transfer activation reactions (Scheme V), the difference between them must also be emphasized since the radical ion pair in Scheme V does not dissociate but disappears by the subsequent proton-transfer process, followed by the next highly exothermic electron-transfer process. Such a situation may be the main factor that has created controversies.[6-19,22,23] Nonetheless, the radical ion pair in Scheme V has a reality in the sense to provide a quantitative description of the energetic profile of the hydride-transfer reactions from PyH_2 to Q as shown in this study. The generality of the electron-transfer activation mechanism as well as the direct detection of the radical ion pair in the endothermic region, which seems to be very difficult by the conventional methods, must be scrutinized by further studies, since the systems in this study are limited in the number owing to its strong dependence on the experimental determinations of various thermodynamic values.

Registry No. $[RuL_3]^{2+}$, 15158-62-0; $[RuL_2L']$, 64189-97-5; 4-MeOBNAH, 105762-82-1; 4-MeBNAH, 19355-10-3; BNAH, 952-92-1; 4-ClBNAH, 105762-83-2; 2,4-Cl_2BNAH, 105762-84-3; BCQH, 17260-79-6; 4-ClBNAH$^{•+}$, 105762-85-4; 2,4-Cl_2BNAH$^{•+}$, 105762-86-5; $AcH_2^{•+}$, 105784-79-0; 4-MeBNA$^+$, 61777-43-3; BNA$^+$, 15519-25-2; 4-ClBNA$^+$, 105762-87-6; 2,4-Cl_2BNA$^+$, 105762-88-7; AcH$^+$, 13367-81-2; Hantzsch's ester, 1149-23-1; *N,N*-dimethylaniline, 121-69-7; *N,N*-dimethyl-*p*-toluidine, 99-97-8; *N,N'*-diphenyl-*p*-phenylenediamine, 74-31-7; diethyl fumarate, 623-91-6; diethyl terephthalate, 636-09-9; benzophenone, 119-61-9; 1-cyanonaphthalene, 86-53-3; acetophenone, 98-

(74) Equation 38 is derived from eq 19 and 37, since $-2.3RT \log K_{et} = \Delta G_{23}$.

(75) The calculated line (Figure 8d) is drawn in the region where the linear relations of $\Delta G°_{H^+}$ and $\Delta G°_{et}$ with $\Delta G°_{et}$ are shown to hold (Figure 5).

(76) The agreements of the calculated values with the experimental values for X-BNAH and AcH_2 in MeCN as well as NADH in H_2O (Figure 8) indicate also that the large work term w_p may be ascribed mainly to the strong interaction in the tight radical ion pair formed via the CT complex. Although the radical ion pair may be strongly solvated, few or no solvent molecules may be involved between the ions. This is the most important difference between outersphere and innersphere electron-transfer reactions (see ref 2a). See, also: Suppan, P. *J. Chem. Soc., Faraday Trans. 1* **1986**, *82*, 509.

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86-2; propiophenone, 93-55-0; *trans*-stilbene, 103-30-0; 4-methylacetophenone, 122-00-9; 4-methoxyacetophenone, 100-06-1; cyanobenzene, 100-47-0; methyl benzoate, 93-58-3; ethyl benzoate, 93-89-0; 3,5-dichloropyridine, 2457-47-8; 3-cyanopyridine, 100-70-9; 4-cyanopyridine, 100-48-1; 3-bromopyridine, 626-55-1; 3-acetylpyridine, 350-03-8; 4-acetylpyridine, 1122-54-9; pyridine, 110-86-1; 2-aminopyridine, 504-29-0; 2,3-dichloro-5,6-dicyano-*p*-benzene, 84-58-2; 2,3-dicyano-*p*-benzoquinone, 4622-04-2; *p*-chloranil, 118-75-2; *p*-bromanil, 488-48-2; 2,6-dichloro-*p*-benzoquinone, 697-91-6; 2,5-dichloro-*p*-benzoquinone, 615-93-0; chloro-*p*-benzoquinone, 695-99-8; *p*-benzoquinone, 106-51-4; methyl-*p*-benzoquinone, 553-97-9; 2,6-dimethyl-*p*-benzoquinone, 527-61-7; trimethyl-*p*-benzoquinone, 935-92-2; magnesium, 22537-22-0.

Isotopic Multiplets in the Carbon-13 NMR Spectra of Aniline Derivatives and Nucleosides with Partially Deuterated Amino Groups: Effects of Intra- and Intermolecular Hydrogen Bonding

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Abstract: In aniline derivatives, the carbon-13 resonances of atoms bearing partially deuterated amino groups, as well as the resonances of vicinal carbon atoms, appear as multiplets. This phenomenon, which is due to upfield deuterium isotope effects on carbon-13 chemical shifts, is observed under conditions of slow hydrogen exchange (e.g., in Me_2SO solutions). The effects are larger for groups engaged in intramolecular hydrogen bonds. Empirical expressions are presented that relate isotope effects with amino proton chemical shifts and hydrogen bond energies. Isotopic multiplets are also observed in the carbon-13 NMR spectra of partially deuterated nucleosides. The multiplet structure is altered upon formation of base pairs. These results are interpreted in terms of hydrogen exchange reactions involving uridine (or thymidine) hydrogen-bonded dimers or changes in hydrogen bond energies upon formation of guanosine–cytidine complexes. Estimates are given for the energies of individual hydrogen bonds in Watson–Crick base pairs.

The conformation, properties, and function of organic molecules in aqueous solution, including biological and synthetic polymers, depend to a large extent on the phenomenon of hydrogen bonding. Because of its fundamental importance, this phenomenon has been the subject of extensive studies by all of the available spectroscopic techniques.[1] Recently a correlation was observed between deuterium isotope effects on carbon-13 chemical shifts and hydrogen bond energies for phenolic and enolic hydroxyl groups engaged in intramolecular hydrogen bonds.[2] This paper presents results on isotopic multiplets in carbon-13 NMR spectra and deuterium isotope effects on carbon-13 chemical shifts for aniline derivatives and nucleosides. In selecting compounds for investigation, particular attention was paid to the possible presence or absence of hydrogen bonding phenomena.

Carbon atoms in the vicinity of partially deuterated functional groups containing exchangeable hydrogens exhibit multiplet structure in the proton-decoupled carbon-13 NMR spectrum. These multiplets result from small upfield deuterium isotope effects on the carbon-13 chemical shifts and are observable under conditions of slow (relative to the magnitude of the isotope effect) chemical exchange between the protio and deuterio forms.[3] Hydrogen ion exchange for amino groups attached to aliphatic carbons is usually fast but slows down considerably upon protonation or coordination to a metal ion. Indeed, isotopic multiplets have been observed in the carbon-13 NMR spectra of ammonium ion derivatives[4] and cobalt(III) complexes.[5] The basicities of aniline derivatives as well as of other conjugated amines are much lower than those of aliphatic amines. Resolved proton resonances of amino and imino groups of such materials in nonhydroxylic solvents can be observed and the results of spin coupling to protons on α carbons can be discerned.[6] Therefore, the observation of isotopic multiplets in their carbon-13 NMR spectra could be anticipated.

Intramolecular hydrogen bonding in aniline derivatives with carbonyl or nitro groups in the ortho position is a well-known phenomenon established on the basis of infrared spectral evidence.[7] The amino proton shifts also indicate the presence of such interactions.[8] Watson–Crick base pairing of nucleosides through intermolecular hydrogen bonds has been the subject of numerous investigations by NMR spectroscopy. Thus, e.g., Katz and Penman showed that formation of stable complexes between guanosine and cytidine in Me_2SO solutions leads to substantial downfield shifts of the amino and imino protons.[9] Subsequently Newmark and Cantor obtained thermodynamic data in the same system from the concentration and temperature dependencies of the proton shifts.[10] More recently Petersen and Led conducted a similar investigation by carbon-13 NMR spectroscopy.[11] Substantial downfield shifts upon base pairing were observed for all of the base carbons, except for carbon-5 of guanosine, which shifted upfield. However, in Me_2SO as the solvent no interaction between adenosine and uridine or thymidine could be detected either by proton[9] or carbon-13 NMR.[12] In order to study the

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